

## Dissociation Constants for Acids at 25°C

Name	Formula	$K_{a1}$	$K_{a2}$	$K_{a3}$
Acetic	HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub>	$1.80 \times 10^{-5}$		
<b>Arsenic</b>	<b>H<sub>3</sub>AsO<sub>4</sub></b>	<b><math>5.60 \times 10^{-3}</math></b>	<b><math>1.00 \times 10^{-7}</math></b>	<b><math>3.00 \times 10^{-12}</math></b>
Arsenous	H <sub>3</sub> AsO <sub>3</sub>	$5.10 \times 10^{-10}$		
<b>Ascorbic</b>	<b>HC<sub>6</sub>H<sub>7</sub>O<sub>6</sub></b>	<b><math>8.00 \times 10^{-5}</math></b>	<b><math>1.60 \times 10^{-12}</math></b>	
Benzoic	HC <sub>7</sub> H <sub>5</sub> O <sub>2</sub>	$6.30 \times 10^{-5}$		
<b>Boric</b>	<b>H<sub>3</sub>BO<sub>3</sub></b>	<b><math>5.80 \times 10^{-10}</math></b>		
Butanoic	HC <sub>4</sub> H <sub>7</sub> O <sub>2</sub>	$1.50 \times 10^{-5}$		
<b>Carbonic</b>	<b>H<sub>2</sub>CO<sub>3</sub></b>	<b><math>4.30 \times 10^{-7}</math></b>	<b><math>5.60 \times 10^{-11}</math></b>	
Chloroacetic	HC <sub>2</sub> H <sub>2</sub> O <sub>2</sub> Cl	$1.40 \times 10^{-3}$		
<b>Chlorous</b>	<b>HClO<sub>2</sub></b>	<b><math>1.10 \times 10^{-2}</math></b>		
Citric	H <sub>3</sub> C <sub>6</sub> H <sub>5</sub> O <sub>7</sub>	$7.40 \times 10^{-4}$	$1.70 \times 10^{-5}$	$4.00 \times 10^{-7}$
<b>Cyanic</b>	<b>HCNO</b>	<b><math>3.50 \times 10^{-4}</math></b>		
Formic	HCNO <sub>2</sub>	$1.80 \times 10^{-4}$		
<b>Hydroazoic</b>	<b>HN<sub>3</sub></b>	<b><math>1.90 \times 10^{-5}</math></b>		
Hydrocyanic	HCN	$4.90 \times 10^{-10}$		
<b>Hydrofluoric</b>	<b>HF</b>	<b><math>6.80 \times 10^{-4}</math></b>		
Hydrogen chromate ion	HCrO <sub>4</sub> <sup>-</sup>	$3.00 \times 10^{-7}$		
<b>Hydrogen peroxide</b>	<b>H<sub>2</sub>O<sub>2</sub></b>	<b><math>2.40 \times 10^{-12}</math></b>		
Hydrogen selenate ion	HSeO <sub>4</sub> <sup>-</sup>	$2.20 \times 10^{-2}$		
<b>Hydrosulfuric</b>	<b>H<sub>2</sub>S</b>	<b><math>9.50 \times 10^{-8}</math></b>	<b><math>1.00 \times 10^{-19}</math></b>	
Hypobromous	HBrO	$2.50 \times 10^{-9}$		
<b>Hypochlorous</b>	<b>HClO</b>	<b><math>3.00 \times 10^{-8}</math></b>		
Hypoiodous	HIO	$2.30 \times 10^{-11}$		
<b>Iodic</b>	<b>HIO<sub>3</sub></b>	<b><math>1.70 \times 10^{-1}</math></b>		
Lactic	HC <sub>3</sub> H <sub>5</sub> O <sub>3</sub>	$1.40 \times 10^{-4}$		
<b>Malonic</b>	<b>H<sub>2</sub>C<sub>3</sub>H<sub>2</sub>O<sub>4</sub></b>	<b><math>1.50 \times 10^{-3}</math></b>	<b><math>2.00 \times 10^{-6}</math></b>	
Nitrous	HNO <sub>2</sub>	$4.50 \times 10^{-4}$		
<b>Oxalic</b>	<b>H<sub>2</sub>C<sub>2</sub>O<sub>4</sub></b>	<b><math>5.90 \times 10^{-2}</math></b>	<b><math>6.40 \times 10^{-5}</math></b>	
Paraperiodic	H <sub>5</sub> IO <sub>6</sub>	$2.80 \times 10^{-2}$	$5.30 \times 10^{-9}$	
<b>Phenol</b>	<b>HC<sub>6</sub>H<sub>5</sub>O</b>	<b><math>1.30 \times 10^{-10}</math></b>		
Phosphoric	H <sub>3</sub> PO <sub>4</sub>	$7.50 \times 10^{-3}$	$6.20 \times 10^{-8}$	$4.20 \times 10^{-13}$
<b>Propionic</b>	<b>HC<sub>3</sub>H<sub>5</sub>O<sub>2</sub></b>	<b><math>1.30 \times 10^{-5}</math></b>		
Pyrophosphoric	H <sub>4</sub> P <sub>2</sub> O <sub>7</sub>	$3.00 \times 10^{-2}$	$4.40 \times 10^{-3}$	
<b>Selenous</b>	<b>H<sub>2</sub>SeO<sub>3</sub></b>	<b><math>2.30 \times 10^{-3}</math></b>	<b><math>5.30 \times 10^{-9}</math></b>	
Sulfuric	H <sub>2</sub> SO <sub>4</sub>	Strong acid	$1.20 \times 10^{-2}$	
<b>Sulfurous</b>	<b>H<sub>2</sub>SO<sub>3</sub></b>	<b><math>1.70 \times 10^{-2}</math></b>	<b><math>6.40 \times 10^{-8}</math></b>	
Tartaric	H <sub>2</sub> C <sub>4</sub> H <sub>4</sub> O <sub>6</sub>	$1.00 \times 10^{-3}$	$4.60 \times 10^{-5}$	